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(54) **Process for pickling and passivating stainless steel without using nitric acid.**

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PATENT ABSTRACTS OF JAPAN, vol. 11, no. 80 (C-409)(2527), 11 March 1987; & JP-A-61 235 581

PATENT ABSTRACTS OF JAPAN, vol. 6, no. 106 (C-108)(984), 16 June 1982; & JP-A-57 035 686

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Description

Technical problem

During the manufacture of hot-rolled steel products or of intermediate products undergoing heat treatment such as annealing, it is well known that the material becomes covered with a layer of oxidation products of varying thickness. Because of the need to obtain a bright surface finish for the final product, these oxide layers have to be totally removed. This is achieved by the well known pickling processes, for which inorganic mineral acids such as hydrochloric, sulphuric, nitric and hydrofluoric acids are used, either alone or in mixtures of varying proportions.

In the stainless steel field, based on a knowledge of currently used industrial processes the pickling process most commonly or indeed almost exclusively used involves the use of a mixture of nitric acid and hydrofluoric acid, the mutual concentrations of which vary according to the type of plant, the type of steel to be pickled, its surface characteristics and the geometry of the article to be treated. The process is without doubt economical and enables excellent results to be obtained. It has however the very serious drawback of creating considerable problems of an ecological nature which are difficult to solve, precisely because of the use of nitric acid. In this respect, nitrogen oxide vapours of general formula NO_x are emitted into the atmosphere, these being highly polluting and aggressive towards metals and non-metals with which they come into contact, and in addition high nitrate levels are attained in the wash water and in the spent baths, with the consequent problem of their disposal. The elimination of the NO_x vapours in the air and of the nitrates in spent baths creates considerable plant problems (for example at the moment there is no NO_x treatment method which is technically free of problems), high running costs and no certainty that the results will satisfy current regulations. Thus in the final analysis the cost in terms of investment is difficult to sustain in most industrial plants.

A pickling system which does not involve the use of nitric acid is therefore of considerable industrial interest and various proposals have been advanced in this respect throughout the world, especially during the last ten years.

Processes alternative to the use of nitric acid: state of the art

A search carried out of patents relating to nitric acid-free cycles proposed as an alternative to the traditional stainless steel pickling process based on $\text{HNO}_3 + \text{HF}$ and of the main technical literature on

this subject has brought to light the following:

A) Japanese patent JP 50071524 published on 13.6.75 uses a system consisting of hydrochloric acid and ferric chloride at a temperature of 70°C and a treatment time of 20 seconds;

B) the two Japanese patents JP 55018552 published on 8.2.80 and JP 55050468 published on 12.4.80 comprise three stages, namely: 1) an initial descaling in sulphuric or hydrochloric acid, 2) subsequent immersion firstly in a solution of potassium permanganate and inorganic acids (not HF) and secondly in a solution of ferric nitrate, ferric sulphate and peroxydisulphuric acid, and 3) final washing with pressurized water or ultrasound;

C) Swedish patent SE 8001911 published on 12.10.81 describes treatment, for a time of between 1 and 120 minutes (1-20 mins preferred) at a temperature of between 10 and 90°C (30 - 60°C preferred), in a solution formed from sulphuric acid and hydrogen peroxide;

D) German patent DD 244362 published on 1.4.87 uses at 15 - 30°C a solution formed from chromic acid, sulphuric acid, hydrofluoric acid and an inhibitor (hexamethylenetetramine); the bath is then neutralized with calcium and barium salts;

E) German patent DE 3937438 published on 30.8.90 is mainly directed towards the wire processing industry and uses a hydrofluoric acid solution containing Fe^{3+} added in the form of a fluoride complex; an oxygenated gas and/or fluid medium is then added to the solution to be subjected to an electrolysis process to obtain nascent oxygen able to oxidize the bivalent iron to trivalent;

F) German patent DE 3222532 published on 22.12.83 describes the pickling of austenitic steel pipes or vessels, the internal surfaces of which are treated at 15 - 30°C with a solution formed from hydrofluoric acid and peroxides (either stabilized hydrogen peroxide or sodium perborate or organic peroxides not further identified), whereas the external surfaces are pickled with pastes formed from hydrofluoric acid, peroxides and fillers (carboxymethylcellulose); the pastes have to be disposed of by neutralization with calcium salts, the peroxides being destroyed either by catalysts or by heating;

G) British patent 2,000,196 of TOKAI Denka Kogyo uses a pickling bath consisting of ferric sulphate and hydrofluoric acid. H_2SO_4 and hydrogen peroxide in a 1:1 molar ratio are fed continuously to maintain an adequate ferric ion concentration during the process. The method for controlling the process by continuously measuring the Redox potential of the system is also claimed, this having to be maintained at ≥ 300

mV by controlling the feed of $\text{H}_2\text{SO}_4 + \text{H}_2\text{O}_2$;
 H) two very similar European patents EP 188975
 and EP 236354 (= WO 87/01739) with priority
 dates 22.1.85 and 19.9.85 respectively, use a
 pickling solution consisting of hydrofluoric acid
 (5-50 g/l) and a trivalent ferric ion introduced in
 the form of fluorinated complexes, into which air
 or oxygen is continuously blown; the treatment
 time is between 30 seconds and 5 minutes and
 the temperature varies from 10°C to 70°C ; it is
 also recommended to continuously control the
 Redox potential, which has to be maintained
 between -200 and +800 mV for the first patent
 and between +100 mV and +300 mV for the
 second patent, adding an oxidant such as potas-
 sium permanganate or hydrogen peroxide to
 raise the potential if necessary. All the tests
 carried out relate only to pickling steel sheet,
 without using oxidizing compounds other than
 atmospheric air.

Finally, there are two further patents relating to
 the possibility of preventing or reducing to a mini-
 mum the formation of NO_x nitrogen oxides in baths
 operating with nitric acid, by directly feeding suit-
 able oxidants into the pickling bath. The first, Japa-
 nese JP 58110682 of 1.7.83, uses hydrogen perox-
 ide; the other, Swedish (SE 8305648 of 15.4.85,
 priority date 14.10.83, SE 835648), again uses hy-
 drogen peroxide and/or alternatively urea.

Notwithstanding this proliferation of patents, un-
 til to-day the traditional process based on nitric
 acid and hydrofluoric acid is still widely used
 throughout the world and none of the aforesaid
 proposed alternatives have been accepted industri-
 ally.

Process of the invention

The process according to the present patent
 application has shown brilliant results both in lab-
 oratory tests and, in particular, in industrial trials
 carried out on high-production lines and plants, and
 is undoubtedly superior to all previous proposals. It
 incorporates interesting aspects of certain of these
 proposals, which are rationalized into an overall
 and exhaustive project to which various aspects of
 absolutely novel character are added.

The process is based on the use of a pickling
 bath containing ferric ions, H_2SO_4 , HF, H_2O_2 and
 conventional additives such as wetting agents,
 emulsifiers, brighteners and anticorrosives, into
 which a strong air flow is continuously blown. The
 operating temperature is generally between 30 and
 70°C and preferably between 45 and 55°C . The
 basic characteristics of the process are as follows:
Inorganic mineral acid content of the bath: on pre-
 paring the pickling bath a solution is prepared
 containing at least 150 g/l and preferably about 170

g/l of H_2SO_4 , and at least 40 g/l and preferably
 about 50 g/l of HF. These acids have various
 functions, of which the most important are to main-
 tain the process pH at less than 1 and preferably
 between 0 and 0.5, to solubilize the oxides originat-
 ing from the heat treatment, and, in the case of the
 hydrofluoric acid, to complex the Fe^{3+} and Cr^{3+}
 ions to the maximum extent.

As the concentrations of the two acids, and
 particularly the hydrofluoric acid, tend to fall during
 the pickling process, they have to be fed periodi-
 cally on the basis of the bath analysis (free acid
 and fluoride ion values).

Fe^{3+} ion content of the bath: on preparing the
 bath, an Fe^{3+} ion quantity of not less than 15 g/l is
 introduced into the pickling solution in the form of
 ferric sulphate. The function of this ion is to replace
 nitric acid as oxidizing agent in the reaction 2Fe^{3+}
 $+ \text{Fe} \rightarrow 3\text{Fe}^{2+}$, favoured by the bath pH con-
 ditions. During the process the correct conditions
 for maximizing the ferric rather than ferrous form
 for the iron dissolved in the bath must be continu-
 ously created.

The oxidation of Fe^{2+} ions to Fe^{3+} ions during
 the process to maintain the concentration of these
 latter above the minimum predetermined value is
 achieved by the combined action of the air blown
 into the bath and the H_2O_2 added continuously to
 the bath in small quantities.

Continuous addition of stabilized hydrogen perox- ide

For the process to be economical, the quantity
 of hydrogen peroxide consumed must be as small
 as possible. For this reason it is important to use
 hydrogen peroxide containing a known stabilizer
 effective in preventing or at least substantially re-
 tarding peroxide decomposition under the operation
 conditions (temperature up to 70°C , very acid bath
 pH, iron up to 100 g/l, presence of Ni and Cr ions).
 A particularly suitable stabilizer is that marketed by
 Interlox (Laporte-Solvay) under the name Interlox S
 333 or Interlox S 333C. The use of suitably stabi-
 lized H_2O_2 in combination with the use of air blown
 into the bath as complementary oxidation medium
 has resulted in a process in which the use of H_2O_2
 is economically convenient, this never having been
 possible with known processes. The pickling bath
 is prepared with an H_2O_2 concentration of between
 1 and 20 g/l, and preferably 2-5 g/l.

During pickling, the continuous H_2O_2 feed is
 regulated on the basis of the type of steel to be
 pickled, the surface characteristics of the material
 (or semi-finished product), and the quantity and
 quality of scale resulting from rolling or annealing.
 Generally an H_2O_2 quantity of between 0.3 and 1
 g/l of bath per hour of operation is fed.

Continuous air blowing

During pickling, a continuous air flow into the bath is maintained at a rate of at least 3 m³/m³ of bath per hour of operation. This air flow, if fed in at a suitable speed, contributes to good bath agitation, this being an important condition for effective pickling as it continuously disturbs the laminar layer in proximity to the surface to be treated, hence ensuring that this surface is always in direct contact with a fresh pickling solution. To ensure optimum mechanical agitation and homogenization of the treatment liquid it is advisable to blow the air into the bottom of the tank via perforated feed tubes, or by the use of suitable blowing equipment.

As already stated, the blown air also oxidises the ferrous ions in combination with the hydrogen peroxide, so resulting in a considerable reduction in the consumption of this latter.

Control of Redox potential: it is well known that the behaviour of stainless steel in acid mixtures is characterised by polarization curves which present activity, passivity and transpassivity phases for different potentials, so that the bath must be kept under those conditions in which the material does not corrode, ie the process must be operated at a potential falling within the passivity range, which can be predetermined based on the type of steel.

During operation, as the concentration of the bivalent ferrous ion in the bath increases, the Redox potential of the bath tends to decrease, however the addition of hydrogen peroxide in combination with the oxidising action of the blown air returns it to optimum values, normally well in excess of 300 mV. By constantly controlling the potential it is therefore possible not only to ensure good material pickling but also to ensure that the passivation film forms on it. In this respect, industrial tests have resulted in all cases in bright, shiny and perfectly level surfaces on which no signs of any corrosive attack due for example to pitting or excessive pickling action were visible. In this respect it should be noted that often the traditional pickling process based on nitric acid and hydrofluoric acid results in such defects, and it is by no means rare for the traditional system to result in "burning" of the material (ie intercrystalline corrosion phenomena).

During those periods in which the pickling bath is not operating (weekend, nights), it is necessary only to provide minimum air blowing to maintain the Redox potential at optimum values, so making it possible to leave the material immersed in the solution for many hours without risk of attack.

Miscellaneous additives content of the pickling bath

In formulating the pickling bath according to the present invention, usual additives for this type of process are used, chosen from non-ionic surfactants acting as wetting agents, emulsifiers, brighteners and acid attack inhibitors. These additives, by mutual synergic action, improve and favour the pickling action. They are used in a total quantity of about 1 g/l of bath.

Advantages of the process

Absence of sludge: the process according to the invention reduces to a minimum or even prevents sludge formation with consequent further cost saving. This advantage is due inter alia to an appropriate HF concentration during the process and to proper control of the concentration of ferrous ions, which are immediately and adequately oxidized to ferric ions.

Facility for automatic control: the process can be constantly controlled by automatic equipment which on the basis of analytical measurements (free and total acid, free fluoride ion content, bivalent ferrous ion content, Redox potential) meter the quantities of pickling products and stabilized hydrogen peroxide to be fed to achieve correct operating parameters.

Process versatility: the process of the invention is easily adaptable to all industrial stainless steel treatment plants, requiring only modest modification. It is also suitable for treating articles and semifinished products of any type, including wire, rod, strip, plate and tubes, the treatment parameters (temperature, time, concentrations) being able to undergo variation without in any way prejudicing results.

The process is suitable for steel of any type: martensitic, ferritic, austenitic.

The following examples are given merely to illustrate some applications of the process according to the invention.

A) Tests on an industrial plant, processing steel rod

70 t of steel rod of average diameter 6 mm, equivalent to about 5000 m² of the following materials: AISI 303, AISI 304 L, AISI 304 K, AISI 304 K2, AISI 316 L, AISI 316 R, AISI 316 Ti and AISI 430, were treated in an industrial tank with a useful bath capacity of 5 m³.

The initial pickling bath had the following composition:

172 g/l of H₂SO₄
48 g/l of HF
15 g/l of Fe³⁺
5 g/l of H₂O₂

2 g/l of H₂O₂ stabilizer

1 g/l of miscellaneous additives.

130 vol. hydrogen peroxide was used. The hydrogen peroxide stabilizer was Interlox S 333 of Laporte Interlox.

The additives consisted of non-ionic surfactants and acid attack inhibitors of known type for pickling baths.

The initially measured Redox potential was about 700 mV.

During the test, which lasted a total of 300 hours, stabilized hydrogen peroxide was added continuously at a rate of 1 g/l per hour of operation. H₂SO₄ was added at intervals to a total of 340 kg, as were HF to a total of 460 kg and additives of the aforesaid type to a total of 25 kg.

The bath temperature was maintained between 50 and 60 °C and the air flow at 30 m³/h.

The treatment time varied between 40 and 75 minutes according to the type of steel treated, with pickling kinetics similar to if not in various cases better than those of the traditional process based on nitric acid and hydrofluoric acid, which was simultaneously compared in a parallel tank.

The Redox potential, measured periodically, remained between 350 and 450 mV, hence ensuring optimum surface finish of the material treated. On termination of treatment the total iron content was about 100 g/l with an Fe³⁺ content of 60 g/l and an Fe²⁺ content of 40 g/l.

In no case and on no material was there any surface pitting or "burning".

On termination of treatment the formation of precipitate in the bath was found to be totally irrelevant and consisted mainly of graphite. No ferrous sulphate crystallization was found. The bath was found to still possess full pickling efficiency.

B) Industrial tests on strip and plate

Tests were carried out on AISI 303, AISI 304 and AISI 316 strip and plate in an industrial plant using the process of the invention and the traditional process for comparison.

a) traditional process:

1st tank: electrolytic pickling with H₂SO₄ - 1 minute at 60-70 °C

2nd tank: electrolytic pickling with HNO₃ - 1 minute at 60-70 °C

3rd tank: pickling with HNO₃ + HF - 1 minute at 70 °C.

b) process of the invention:

1st tank: electrolytic pickling with H₂SO₄ - 1 minute at 60-70 °C

2nd tank: treatment for 1 minute at 55-60 °C with the following bath:

150 g/l of H₂SO₄

48 g/l of HF

15 g/l of Fe³⁺

5 g/l of H₂O₂

2 g/l of H₂O₂ stabilizer (Interlox S 333 C)

1 g/l of miscellaneous additives (of the type indicated in the preceding example).

3rd tank: treatment for 1 minute at 55-60 °C with the same bath composition as the 2nd tank.

The useful bath capacity of the 2nd and 3rd tank was each 10,000 litres.

During the test (lasting about 240 hours) 0.6 g/l of H₂O₂ per hour (stabilized as stated) was fed continuously into the bath. No further additions of other ingredients were made. The air flow was maintained at 30 m³/h to each tank. The total material treated in the test was 1800 t.

The surface appearance of the plate on termination of the process was always shiny and bright, and was better than that obtained with the traditional comparison test. There was no evidence of excess pickling or surface pitting on any material.

C) Laboratory tests on tubes

Laboratory tests were carried out on AISI 304 and AISI 316 tubes under the bath conditions described under point A.

The ratio of the material quantity used to the test tank capacity was equal to that of normal industrial cycles. The temperature was fixed at 50 °C and the treatment time varied from 30 to 60 minutes depending on the type of material.

The progress of the test and the results obtained were similar to those described under point A, with regard to product consumption, to the behaviour of the Redox potential, to the final surface appearance of the material, to the attack kinetics and to the absence of any pitting phenomena.

CONCLUSIONS of the industrial scale trials.

From the foregoing it is apparent that the new stainless steel pickling and passivation process, characterised by a bath of specific composition, control of the bath during the entire operation, in particular of its Redox potential, and continuous air blowing, represents an optimum solution in terms of the technical result of the treatment, process economy (in particular due to the low H₂O₂ consumption), and the pollution problem posed by traditional nitric acid processes.

Claims

1. A process for pickling and passivating stainless steel, consisting of bringing the material to be treated into contact with a bath maintained at a temperature of between 30 and 70 °C and preferably between 45 and 55 °C and having the following initial composition:
 - a) H₂SO₄ at least 150 g/l
 - b) Fe³⁺ at least 15 g/l
 - c) HF at least 40 g/l
 - d) H₂O₂ (containing known stabilizers) 1-20 and preferably 2-5 g/l
 - e) additives of the non-ionic surfactant type (emulsifiers, wetting agents, brighteners) and acid attack inhibitor type: about 1 g/l in total; into said bath there being continuously fed:
 - an air flow of at least 3 m³/h per m³ of bath, using a suitable distributor device for diffusing the flow into the liquid mass;
 - a quantity of stabilized H₂O₂ of between 0.3 and 1 g/l per hour, controlled on the basis of the Redox potential of the bath, which must be maintained at ≥ 350 mV;
 - and possibly sufficient quantities of ingredients a), c) and e) to maintain their concentration in the bath at optimum levels and the bath pH at less than 1 and preferably between 0 and 0.5.
2. A process as claimed in claim 1, wherein the Fe³⁺ ions are introduced into the initial bath in the form of ferric sulphate.
3. A process as claimed in claim 1, wherein H₂O₂ stabilized with the stabilizer interox is used.
4. A process as claimed in claim 1, wherein a bath is used of initial composition:

172 g/l of H₂SO₄
 48 g/l of HF
 15 g/l of Fe³⁺
 5 g/l of H₂O₂
 1 g/l of non-ionic surfactant and acid attack inhibitor additives
 2 g/l of H₂O₂ stabilizer in the form of Interlox S 333.
5. A process as claimed in claim 1, conducted in combination with preliminary partial removal of oxides by a known process.

Patentansprüche

1. Verfahren zum Beizen und Passivieren von rostfreiem Stahl, bestehend aus dem Inkontaktbringen des zu behandelnden Materials mit einem Bad, das bei einer Temperatur von zwischen 30 und 70 °C und vorzugsweise zwischen 45 und 55 °C gehalten wird, und das die folgende Anfangszusammensetzung hat:
 - a) H₂SO₄ zumindest 150 g/l
 - b) Fe³⁺ zumindest 15 g/l
 - c) HF zumindest 40 g/l
 - d) H₂O₂ (umfassend bekannte Stabilisatoren) 1-20 und vorzugsweise 2 bis 5 g/l
 - e) Additive vom Typ des nicht ionischen oberflächenaktiven Mittels (Emulgatoren, Benetzungsmittel, Aufhellmittel) und vom Typ des Säureangriffsinhibitors: etwa 1 g/l insgesamt; wobei in das Bad kontinuierlich geführt werden:
 - ein Luftfluß von wenigstens 3 m³/h pro m³ Bad, unter Verwendung eines geeigneten Verteilungsgerätes zum diffundieren des Flusses in die flüssige Masse;
 - eine Menge an stabilisiertem H₂O₂ von zwischen 0,3 bis 1 g/l pro Stunde, gesteuert auf der Basis des Redoxpotentials des Bads, das bei ≥ 350 mV gehalten werden muß;
 - und gegebenenfalls ausreichende Mengen an Bestandteilen (a), c) und e), um deren Konzentration in dem Bad bei optimalen Gehalten und den Bad-pH bei weniger als 1 und vorzugsweise zwischen 0 und 0,5 zu halten.
2. Verfahren nach Anspruch 1, dadurch **gekennzeichnet**, daß die Fe³⁺ -Ionen in das Anfangsbad in der Form von Ferrisulphat eingeführt werden.
3. Verfahren nach Anspruch 1, dadurch **gekennzeichnet**, daß H₂O₂ verwendet wird, das mit dem Stabilisator Interlox stabilisiert ist.
4. Verfahren nach Anspruch 1, dadurch **gekennzeichnet**, daß ein Bad der Anfangszusammensetzung verwendet wird:

172 g/l H₂SO₄
 48 g/l HF
 15 g/l Fe³⁺
 5 g/l H₂O₂
 1 g/l an nicht ionischem oberflächenakti-

vern Mittel und Säureangriffsinhibitor-Additven
2 g/l H₂O₂ - Stabilisator in der Form von
Interlox S 333.

inhibiteurs d'attaque acide
2 g/l de stabilisant d'H₂O₂ sous forme
d'Interlox S 333.

5. Verfahren nach Anspruch 1,
durchgeführt in Kombination mit der vorherigen
teilweisen Entfernung von Oxiden durch
ein bekanntes Verfahren.

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5. Procédé selon la revendication 1, réalisé en
combinaison avec l'élimination partielle préalable
des oxydes par une procédé connu.

Revendications

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1. Procédé de décapage et de passivation d'acier
inoxydable consistant à porter le matériau à
traiter au contact d'un bain maintenu à une
température comprise entre 30 et 70 °C, et de
préférence entre 45 et 55 °C, et ayant la composition
initiale suivante :

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- a) H₂SO₄ au moins 150 g/l
- b) Fe³⁺ au moins 15 g/l
- c) HF au moins 40 g/l
- d) H₂O₂ (contenant des stabilisants connus)
1-20, et de préférence 2-5 g/l.
- e) additifs du type surfactif non ionique
(émulsifiants, agents mouillants, lustrants) et
du type inhibiteur d'attaque acide :

20

environ 1 g/l au total; en introduisant de
manière continue dans ledit bain :

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- un courant d'air d'au moins 3 m³/h par
m³ de bain, en utilisant un dispositif
distributeur approprié pour diffuser le
courant dans la masse liquide;
- une quantité d'H₂O₂ stabilisé comprise
entre 0,3 et 1 g/l par heure, régulée
sur la base du potentiel redox du
bain, qui doit être maintenu à ≥ 350
mV;
- et le cas échéant des quantités suffisantes
des ingrédients a), c) et e)
pour maintenir leur concentration dans
le bain à un niveau optimum, et le pH
du bain à moins de 1, et de préférence
entre 0 et 0,5.

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2. Procédé selon la revendication 1, dans lequel
les ions Fe³⁺ sont introduits dans le bain initial
sous forme de sulfate ferrique.

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3. Procédé selon la revendication 1, dans lequel
on utilise H₂O₂ stabilisé par le stabilisant Interlox.

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4. Procédé selon la revendication 1, dans lequel
on utilise un bain ayant la composition initiale :

172 g/l de H₂SO₄

48 g/l de HF

15 g/l de Fe³⁺

5 g/l de H₂O₂

1 g/l de surfactif non ionique et d'additifs

55