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(54) CATALYST FOR CATALYTIC CRACKING OF HYDROCARBON, WHICH IS USED IN PRODUCTION OF LIGHT OLEFIN AND PRODUCTION METHOD THEREOF

KATALYSATOR ZUM KATALYTISCHEN CRACKEN VON KOHLENWASSERSTOFF, DER BEI DER PRODUKTION VON LEICHTEM OLEFIN VERWENDET WIRD, UND HERSTELLUNGSVERFAHREN DAFÜR

CATALYSEUR POUR LE CRAQUAGE CATALYTIQUE D'UN HYDROCARBURE, UTILISÉ DANS LA PRODUCTION D'UNE OLÉFINE LÉGÈRE, ET PROCÉDÉ DE PRODUCTION DE CE CATALYSEUR

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(73) Proprietor: **SK Innovation Co., Ltd.**

Seoul 110-110 (KR)

(72) Inventors:

- **CHOO, Dae Hyun**
Busan 606-080 (KR)
- **KIM, Hong Chan**
Daejeon 305-712 (KR)
- **KIM, Suk Joon**
Daejeon 305-761 (KR)
- **KIM, Ji Min**
Daejeon 305-761 (KR)
- **KIM, Tae Jin**
Daejeon 305-712 (KR)
- **CHOI, Sun**
Daejeon, 305-712 (KR)
- **OH, Seung Hoon**
Seoul 135-280 (KR)

• **KIM, Yong Seung**

Daejeon 305-712 (KR)

• **PARK, Deuk Soo**

Yuseong-gu

Daejeon 305-712 (KR)

• **PARK, Young Ki**

Daejeon 305-755 (KR)

• **LEE, Chul Wee**

Daejeon 305-150 (KR)

• **KIM, Hee Young**

Daejeon 305-755 (KR)

• **CHOI, Won Choon**

Daejeon, 305-712 (KR)

• **KANG, Na Young**

Yesan-gun

Chungcheongnam-do 340-932 (KR)

• **SONG, Bu Sub**

Daejeon 306-120 (KR)

(74) Representative: **Hubert, Philippe et al**

Cabinet Beau de Loménie

158, rue de l'Université

75340 Paris Cedex 07 (FR)

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Description

[Technical Field]

[0001] The present invention relates to a hydrothermally stable porous molecular sieve catalyst and a preparation method thereof, and more particularly to a hydrothermally stable porous molecular sieve catalyst, which, even in an atmosphere of high temperature and humidity, has a relatively stable structure and can maintain its catalytic activity, and a method of preparing the same.

[Background Art]

[0002] Porous inorganic materials having a framework of -Si-OH-Al- groups have been widely used in the field of porous molecular sieve catalysts because they have abundant pores, large specific surface area, and many active sites and acid sites.

[0003] This porous molecular sieve catalyst is used in, for example, heterogeneous catalytic reactions, such as various oxidation/reduction reactions, including catalytic cracking reactions, isomerization reactions and esterification reactions, particularly heterogeneous catalytic reactions requiring thermal stability under a severe atmosphere of high temperature and humidity. In this case, however, the catalyst has problems in that, when it is placed in a steam atmosphere of more than 500°C, dealumination of its tetrahedral framework will occur, leading to its structural breakdown, and at the same time, the acid sites of the catalyst will be reduced, resulting in a rapid reduction in catalytic activity. Additionally, since high mechanical strength is required for these microporous molecular sieve catalysts in order to be used in massive fluidized catalytic petrochemical processes for naphtha catalytic cracking, inorganic complex and matrix (clay) are used for producing spherical catalysts in these area.

[0004] Therefore, since microporous molecular sieve catalyst comprising many components such as bonding agent, matrix, and porous molecules, maintaining thermal-stability for the respective component is one of the most important factors to produce proper microporous molecular sieve catalyst. For example, the collapse of matrix structure, which is used for microporous molecular sieve catalyst, drastically decreases the reaction rate of naphtha catalytic cracking.

[0005] In other hand, in order to achieve high yield of ethylene and propylene in naphtha catalytic cracking process, it is required to control the characteristic of acid site in microporous molecular zeolite. If the amount of acid site is large or strength of acidity is strong relatively, dehydrogenation reaction is faster, and so the yield of saturated hydrocarbons such as methane and aromatics such as benzene, toluene and xylene, increases.

[0006] On the other hand, if the amount of acid site is small or strength of acidity is weak relatively, conversion of hydrocarbon decreases and so the yield of light olefins decreases.

[0007] As mentioned above, in order to produce light olefins from hydrocarbons such as naphtha by catalytic cracking using catalyst in an effective way, many characteristics of catalyst are required. Specially, the thermal-stability is considered to be the most important factor because the catalytic cracking catalyst is operated in conditions of high temperature and high humidity. Many researches have been proposed to increase thermal-stability.

[0008] Regarding these methods, US patent No. 5,039,644 discloses a method for preparing a catalyst which is stable at high temperature and which comprises 0.5-15 wt% of P_2O_5 embedded in porous metal oxides such as TiO_2 , ZrO_2 , TiO_2 - ZrO_2 mixtures, TiO_2 - Al_2O_3 mixtures, or ZrO_2 - Al_2O_3 mixtures. However, this patent does not explain how to achieve high yield of light olefins from catalytically cracking hydrocarbons using zeolite.

[0009] US patent No. 4,977,122 discloses a hydrothermally stable catalyst, which comprises: (a) a crystalline zeolite; (b) an inorganic oxide matrix (e.g., silica, alumina, silica-alumina, magnesia, zirconia, titania, boria, chromia, clay, etc.); and (c) discrete particles of phosphorus-containing alumina also dispersed in said matrix, said discrete particles having been prepared by contacting alumina with a phosphorus compound selected from the group consisting of an alkaline earth metal salt (Be, Mg, Ca, Sr, Ba) of phosphoric acid or phosphorous acid and mixtures thereof.

[0010] US patent No. 6,835,863 discloses a process for producing light olefins by catalytically cracking naphtha (boiling point: 27-221°C) using a pelletized catalyst containing 5-75% by weight of ZSM-5 and/or ZSM-11, 25-95% by weight of silica or kaolin and 0.5-10% by weight of phosphorus. However, there is no mention of the specific phosphorus starting material or of the hydrothermal stability of the molded catalyst.

[0011] Meanwhile, US patent No. 6,211,104 discloses a catalyst for catalytic cracking, which comprises 10-70 wt% of clay, 5-85 wt% of inorganic oxides and 1-50 wt% of zeolite. The zeolite used in the catalyst consists of 0-25 wt% of Y-zeolite or REY-zeolite, and 75-100 wt% of pentasil zeolite (with $SiO_2/Al_2O_3 = 15-60$) selected from ZSM-5, ZSM-8 and ZSM-11 zeolites containing 2-8 wt% of P_2O_5 and 0.3-3 wt% of Al_2O_3 or MgO or CaO. The starting materials of said aluminum or magnesium or calcium compounds are selected from aqueous solutions of their nitrates, hydrochloride, or sulfates. Particularly, the catalyst is described as showing excellent olefin production even when pretreated in an atmosphere of 100% steam at 800°C for 4-27 hours. However, in said patent, technology for adjusting/selecting and loading the specific chemical species of P is not disclosed, the added metals are limited to Al, Mg and Ca, and a

conventional water-soluble metal salt is used so that the Al, Mg or Ca cations, which are generated during the preparation of the catalyst, can be easily ion-exchanged with the protons of zeolite, resulting in the loss of acidic sites. For this reason, it is believed that it is not easy to prepare the catalyst proposed in said patent under the specified synthesis conditions.

[0012] In US publication No. 2005/0020867 A1, the catalyst for light olefin production is disclosed. Said catalyst is prepared by the steps comprising that ZSM-5 treated with P_2O_5 1~10 wt.%, RE_2O_3 (RE represents rare earth elements) 0~10 wt.%, transition metal (Fe, Co, Ni, Cu, Zn, Mo, Mn) oxides 0.7~15 wt.% is completed by drying and calcination, and then mixed with clay and inorganic bonding agents (silica, alumina, silica-alumina), followed by spray drying. The present ZSM-5 is silica-rich (higher Si/Al ratio) that may reduce aromatization and hydrogen transfer reaction. However, the silica-rich ZSM-5 is not economic because of its complicated synthetic method, weak for matrix performance and structural stability by severe thermal treating with inorganic bonding agents and clay, which are not stable at high temperature steaming. It may cause reducing catalytic cracking activity of zeolite.

[0013] In the US Patent No. 6,613,710, P-modified clay 40~80 wt.%, semi-basic alumina 1~20 wt.%, and ZSM-5 0.5~15 wt.% are used for the catalyst of catalytic cracking reaction. P-modified clay is obtained by treating clay and phosphoric acid at 15~40°C for 1~16 hours, and preparing semi-basic alumina from slurry of sodium aluminate and aluminum sulfate at pH 7.5~9. The present catalyst yields more LPG in residual oil cracking within b.p. 315~528°C. This patent is not for host catalyst but for additive catalyst technology of LPG booster, and there is no disclosure of hydrothermal stabilization improvement and production of light olefins.

[0014] In US patent No. 5,670,037, ZSM-5 modified with rare earth metal and calcined by aluminum phosphate sol is proposed for hydrocarbon catalytic cracking to increase light olefin yield. It is prepared by mixing of P_2O_5 and zeolite (wt. ratio of P_2O_5 to zeolite is 1:5~99) in aluminum phosphate solution, drying, calcining, and steaming. The completed catalyst is made of zeolite 10~35 wt.%, inorganic oxides (Al_2O_3 , SiO_2 , $Al_2O_3-SiO_2$) 5~90 wt.%, and clay 0~70 wt.%. Aluminum phosphate solution is used for treating zeolite, and there is no explanation of the yield increment of light olefins without the usage of rare earth metal.

[0015] In the US Patent No. 6,080,698, the pentasil-type zeolite catalyst for production of light olefin by hydrocarbon catalytic cracking is prepared by treating ZSM-5 ($SiO_2/Al_2O_3 = 15\sim60$) with P_2O_5 1~10 wt.%, alkaline earth metal oxides 0.3~5 wt.%, and transition metal oxides 0.3~5 wt.%. The results with Mg, Ni, Zn, Cu, and Ca for treatment of zeolite are reported, while the result with manganese oxide is not explained. The phosphorus is limitedly used to only modify zeolite with transition metal.

[0016] In the US Patent No. 6,080,303, the zeolite catalyst for production of light olefin by hydrocarbon catalytic cracking is prepared by treating with aluminum phosphate ($AlPO_4$). The catalyst is prepared by 1) making and calcining ZSM-5 modified with phosphorus, 2) forming $AlPO_4$ by mixing $Al(NO_3)_3$ and $NH_4(H_2PO_4)$ at pH 7~9.3) treating phosphorus based ZSM-5 with $AlPO_4$ and calcining. For treatment using $AlPO_4$, both dried state and wet gel state for $AlPO_4$ may be possible. The completed catalyst has a composition comprising P 0.5~10 wt.%, $AlPO_4$ 1~50 wt.%, zeolite 5~60 wt.%, and balanced binder or clay. In the present patent, P and $AlPO_4$ are used to improve hydrothermal stabilization of zeolite, and the advantage of the result of hydrothermal treatment of n-hexane is explained. However, there is no result before hydrothermal treatment, and no explanation of the stabilization technology of binder and clay, as P and $AlPO_4$ are only used for treating zeolite.

[0017] In US Patent 2006/0011513 A1, the catalyst made of ZSM-5, zeolite Beta, Mordenite, Ferrierite, or other zeolite (silica/alumina > 12), which is treated with the mixed binder of aluminum phosphate salts and metal phosphate salts, is proposed as an additive in FCC process. The metal phosphate salts as binder are selected from IIA group, lanthanoids group, Sc, Y, La, Fe, and Ca, and the content of phosphate is more than 5 wt.%, and 4~50 wt.% is included in typical cases. In this patent, chemical structures of phosphate salts, which are not for active sites but for binders, are not shown. Furthermore, there is also no disclosure of the improvement of olefin yield by using zeolite formed with manganese.

[0018] In the US patent 5,380,690, catalyst which comprises clay 0~70%, inorganic oxides such as Al_2O_3 , SiO_2 , $Al_2O_3-SiO_2$ 5~99%, and zeolite 1~50% is disclosed. Said zeolite comprises Y zeolite 0~25%, and P_2O_5 containing ZSM-5 75~100%. Said catalyst is prepared by uniformly mixing ZSM-5 modified from RE_2O_3 1~30%, with aluminum phosphate solution ($Al_2O_3:P_2O_5 = 1:1\sim3$, wt. ratio, P_2O_5 : zeolite=1:5~99), calcining, and steaming.

[0019] The US patent 2006/0116544 reports that by treating pentasil type zeolite - containing rare earth metal, and manganese or zirconium - with phosphorus, hydrothermal stability and yield of light olefin are improved. It is required that manganese or zirconium is included together with rare earth metal and phosphorus in zeolite in order to improve the yield of light olefin. Furthermore, direct injection of rare earth metal and manganese or zirconium and phosphorus in zeolite is used as treating method. The purpose of this technology is structural improvement like the previous ones, and there are no comments about stabilization of inorganic binders or matrix contents.

[0020] In the US patent 4,956,075, the Y zeolite catalyst treated with manganese and rare earth metal is proposed for hydrocarbon catalytic cracking for gasoline with higher octane number. However, the catalyst has less yield of light olefins and hydrothermal stability than pentasil type catalysts.

[0021] Addition of manganese to ZSM-5 may improve hydrothermal stability, reporting in "Studies in Surface Science

and Catalysis", V105, 1549(1996). However, there is only explanation of hydrothermal stability, no explanation for production of light olefins by hydrocarbon catalytic cracking.

[0022] In the US Patent 6,447,741, aluminophosphate treated by manganese is used as a catalyst for catalytic cracking, while there are no results of synthesis of catalyst and application for hydrocarbon cracking. In addition, in this patent,

[0023] US 2007/0209969 discloses a catalyst for catalytic cracking and a method for producing thereof, wherein 6% of P2O4 and 3% of MnO2 are supported as a catalyst component. The catalyst comprises 17% of an alumina binder, 43% of kaolin, and 40% of HZSM-5 and USY zeolites. The phosphorous component and the metal component are supported;

[0024] EP 1762299 discloses a process for preparing a catalyst for cracking hydrocarbons, comprising MPI and HY zeolite, halloysite and montmorillonite clay, alumina inorganic oxide, phosphorus and manganese.

[0025] US 2004/266608 discloses the preparation of a molecular sieve-containing catalyst for cracking hydrocarbons, comprising HY-zeolite, kaolin clay, alumina inorganic oxide, phosphorus and manganese.

[0025] As explained above, transition metals such as manganese, phosphate and rare earth metals have been proposed to increase thermal-stability of catalysts and high yield of light olefins from hydrocarbon catalytic-cracking. However, there is no previous report which explains systematically how to prepare the catalysts for high thermal-stability and high yield of light olefins. That is, there is no previous report as proposed by the present invention, which describes embedding acid site of zeolite by manganese, stabilizing inorganic complex and matrix by phosphate and manganese in order to maintain the catalyst activity for long period and increase yield of light olefins. Also, this present invention shows cost-effective procedure for manufacturing catalyst by eliminating complex embedding step and complex processing spherical catalyst.

[0026] As described in above comparative patents, phosphate shows high ability to increase thermal-stability of zeolite catalyst. Phosphate increases thermal-stability by stabilizing Al through acting as phosphate ion ($[\text{PO}_4]^{3-}$) in -Si-OH-Al-frame which is Bronsted acid site and dealuminated by steam.

[0027] However, thermal-stability is affected strongly by how to introduce phosphate into zeolite. In order to introduce phosphate into zeolite to increase thermal-stability, previous methods tried to inject phosphoric acid directly into zeolite. However, large amount of acid sites are lost according to these methods. Another method is to use phosphoric acid and rare-earth metals, such as La, together. In this method, large size of La^{3+} or phosphoric acid decreases the reaction activity by positioning at entrance of zeolite pore. Additionally since the previous methods tries to make only zeolite itself thermally stable, the problem is that the microporous molecular sieve catalyst made by the zeolite does not have sufficient thermal-stability.

[0028] Therefore, the present intention discloses 1) a method to stabilize the catalyst for long period in circumstances of high temperature and high humidity, 2) a method to maximize yield of light olefins by maintaining acid sites of catalyst after embedding.

[Disclosure]

[Technical Problem]

[0029] The present invention provides a cracking catalyst using components to stabilize the inorganic oxide binder and matrix component added to obtain mechanical strength along with maintaining the structure of zeolite, which is a main catalyst component, under high temperature and high humidity for preparing the cracking catalyst with thermal-stability.

[0030] An aspect of the present invention provides a method of preparing the catalyst, which is easy for mass production and economical due to simple synthesis process, unlike the existing method of preparing the catalyst.

[Technical Solution]

[0031] A hydrocarbon cracking catalyst for preparing light olefin from C4 or more than C4 hydrocarbon having a boiling point of 30 ~ 200 °C, which is characterized in that 0.01~5.0 wt% of MnO_2 and 1~15 wt% of P_2O_5 are simultaneously supported on a catalyst component, wherein the catalyst component comprises 1 ~ 50wt% of zeolite, 21 ~ 70wt% of clay, and 1 ~ 40wt% of an inorganic oxide, wherein zeolite is ZSM-5 zeolite, and wherein both MnO_2 and P_2O_5 are embedded on each of the zeolite, the clay and the inorganic oxide in the catalyst component.

[0032] According to a particular feature, said zeolite has a Si/Al molar ratio of less than or equal to 200.

[0033] A method of preparing the cracking catalyst for preparing light olefin from C4 or more than C4 hydrocarbon, the method comprising the steps of:

(a) Mixing zeolite, clay and inorganic oxide precursor with phosphorus precursor and manganese precursor with

stirring to prepare a mixing slurry; and
(b) spray drying the mixing slurry, followed by calcinations.

[Advantageous Effects]

[0034] The present invention not only improves thermal-stability of the catalyst by embedding manganese and phosphorus in the catalyst comprising zeolite, inorganic oxide and clay simultaneously, but also obtains high yield of light olefin in catalytically cracking hydrocarbons more than C₄ such as naphtha by protecting acid-site of zeolite. Due to simple method of preparing the catalyst, it is easy and economical for mass production.

[Description of Drawings]

[0035] FIG 1 is a schematic diagram of manufacturing the catalyst of this present invention.

[Best Mode]

[0036] Compared to the previous inventions, the present intention discloses a new method to achieve simpler catalyst manufacturing, high thermal-stability, and high yield of light olefins in the area of producing light olefins from cracking hydrocarbons catalytically.

[0037] The method of manufacturing the catalyst of the present inventions is followed:

1. For preparing microporous molecular sieve catalyst, the maximal protection of acid sites of zeolite are accomplished by embedding manganese salt into zeolite in the step of processing slurry of microporous molecular sieve catalyst and by not using the zeolite which previously is embedded by manganese before slurry processing.

2. In order to increase mechanical strength of microporous molecular sieve catalyst, the present invention stabilizes inorganic complex by injecting proper amount of phosphorus and manganese contents in the step of processing slurry of inorganic complex.

3. By mixing zeolite slurry, inorganic complex slurry and clay in final, manganese and phosphorus contents could be embedded in clay, zeolite and inorganic oxide simultaneously so that the stability and activity of decomposition is accomplished to be maximized.

[0038] As described above, although it is well known that injection of phosphate or transition metals into zeolite to stabilize the catalyst structure, the present invention firstly disclosed the effective method to stabilize inorganic complex and clay, and maintain the acid sites of zeolite maximally by embedding simultaneously manganese and phosphorus into zeolite in the step of processing slurry, not in the previous steps of processing zeolite directly, in order to obtain high yield of light olefins in catalytically cracking hydrocarbons more than C₄.

[0039] The catalyst disclosed by the present invention for producing light olefins from hydrocarbons more than C₄ is manufactured by embedding simultaneously 0.01~5.0 wt% of MnO₂ and 1~15 wt% of P₂O₅ into catalyst components which comprise 1~50 wt% zeolite, 21~70 wt% clay, and 1~40 wt% inorganic oxide.

[0040] The catalytic cracking catalyst described above is prepared by the following steps: (a) making mixed slurry by mixing phosphate precursor and manganese precursor into zeolite, clay and inorganic oxide precursor; (b) calcining the above mixed slurry after spray-drying.

[0041] In the examples of this present invention, the mixed slurry, in which phosphate precursor and manganese precursor are mixed into zeolite, clay and inorganic oxide precursor, is prepared, as illustrated in FIG. 1, by steps comprising (i) manufacturing slurry of zeolite and clay by adding and mixing clay after mixing zeolite and manganese precursor; (ii) manufacturing inorganic oxide slurry by mixing phosphate precursor and manganese precursor into inorganic oxide precursor; and (iii) mixing uniformly the above zeolite/clay slurry and inorganic oxide slurry.

[0042] In another examples prepared by this present invention, the mixed slurry, in which phosphate precursor and manganese precursor are mixed into zeolite, clay and inorganic oxide precursor, is prepared by steps comprising (i) manufacturing zeolite slurry by mixing zeolite and manganese precursor; (ii) manufacturing inorganic oxide slurry by mixing phosphate precursor and manganese precursor into inorganic oxide precursor; and (iii) mixing uniformly the above zeolite slurry, clay slurry and inorganic oxide slurry.

[0043] In another examples prepared by the present invention, the mixed slurry, in which phosphate precursor and manganese precursor are mixed into zeolite, clay and inorganic oxide precursor, is prepared by simultaneously mixing zeolite, clay, inorganic oxide precursor, phosphate precursor and manganese precursor.

[0044] Finally, after spray-drying the above mixed slurry, the catalyst for catalytic cracking by the present invention is prepared by calcining 5~10 hours in 500~700°C.

[0045] The catalyst prepared by this method has not only improved hydrothermal stability but higher light olefin yield

in hydrocarbon catalytic cracking, protecting acid sites in zeolite. Activity cannot be guaranteed if the ratio of each component of manganese, phosphorus, zeolite and inorganic oxides in slurry formation for spray drying and mixing progress are not proper.

[0046] Zeolite is ZSM-5 ($\text{Si}/\text{Al} < 200$, mole base).

[0047] Zeolite with $\text{Si}/\text{Al} > 200$ may reduce activity by little acid sites, and the synthesis for such zeolite is not economical. Following the present research, the quantity of zeolite used is 1~50 wt % based on whole catalyst weight.

[0048] Manganese precursor in this invention could be sulfate, nitrate, chloride, or acetate of manganese, and preferable precursors are chloride and acetate of manganese.

[0049] Improvement of light olefin yield is achieved by protecting acid sites of zeolite possibly through agitating with manganese precursor in slurry mixture preparation step of zeolite, clay and inorganic oxides, or zeolite slurry preparation step.

[0050] In the invention manganese precursor is used so that the MnO_2 is 0.01~5.0 wt% based on the final catalyst weight. In case that MnO_2 is less than 0.01 wt%, the protection of acid center and hydrothermal stability decrease. In case that MnO_2 is higher than 5.0 wt%, acid center sharply decreases to lower the activity of catalyst.

[0051] For current invention, clay is used in the range of 21~70 wt% based on the final catalyst weight. In case that the amount of clay is less than 21 wt%, there are many problems of controlling the physical properties such as wear strength and specific gravity. In case that the amount of clay is higher than 70 wt%, catalyst activity could be decreased.

[0052] For the present invention, Al_2O_3 , SiO_2 or $\text{Al}_2\text{O}_3\text{-SiO}_2$ could be used as the binder of inorganic oxidized precursor. For inorganic oxide precursor of catalytic cracking catalyst in current invention, inorganic oxide precursor has the form of sol, gel or solution including Al_2O_3 , SiO_2 , or $\text{Al}_2\text{O}_3\text{-SiO}_3$. The desirable amount of the inorganic oxide is in the range of 1 ~ 40wt% based on the final catalyst weight. When the amount of the inorganic oxide is less than 1 wt%, the wear strength of micro-spherical catalyst could be insufficient, whereas in the case that the amount of inorganic oxidized substance is higher than 40wt%, the activity of catalytic cracking catalyst decreases

[0053] For phosphorus precursor of the present invention, it can be used of the aqueous compound which is selected from the group of H_3PO_4 , $(\text{NH}_4)_3\text{PO}_4$, $\text{H}(\text{NH}_4)_2(\text{PO}_4)$ and $\text{H}_2(\text{NH}_4\text{PO}_4)$, and it is desirable for its contents to have P_2O_5 content of the final catalyst weight to be in the range of 1~15wt%. In the case that P_2O_5 content of the final catalyst is less than 1 wt%, the hydrothermal stability of zeolite decreases, whereas in the case that P_2O_5 content of the final catalyst is higher than 15 wt%, the activity of catalytic cracking decreases due to the excess loss of acid center.

[0054] Phosphorus and manganese contained in mixed slurry are in the dissolved form, embedded to all of zeolite, clay and inorganic oxidized substance. These components protect the acid center of zeolite and increase the hydrothermal stability of zeolite, clay and inorganic oxidized substance to maximize the stability and activity of catalyst.

[0055] Finally, the catalyst for catalytic cracking according to the present invention is prepared by spray-drying and calcining above mixed slurry at 500~700°C for 5~10 hours.

[0056] The prepared catalyst according to the present invention is used as microspheroidal molded catalyst for fluidized catalytic process producing ethylene and propylene from hydrocarbons (carbon number is 4 or above) with high yield and high selectivity, wherein said hydrocarbons (carbon number is 4 or above) mean hydrocarbons which has boiling point of 30~200°C.

[0057] Also, even in the condition of high humidity and high temperature, the catalyst according to the present invention has high cracking activity and stability. Due to this feature, the present catalyst can be used for not only catalytic cracking reaction but also isomerization reaction, alkylation reaction, esterification reaction and oxidation/reduction reaction which require the high hydrothermal stability.

[Mode for Invention]

[0058] Hereinafter, the present invention will be described in more detail using Examples. It is to be understood, however, that these examples are not to be construed to limit the scope of the present invention

Comparative Example 1: Preparation of P-La-Mn/ZSM-5

[0059] 40.5g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ was dissolved in 3000 mL of distilled water. To the solution, 200g of ZSM-5 was slowly added with stirring for about 3 hours at the room temperature. Next, the solution was dried with vacuum drying, followed by being calcined (650°C, 6 hours). 89g of $\text{La}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$ was dissolved in 3000mL of distilled water, and 200g of calcined sample was added to the solution, followed by stirring 3 hours at the room temperature. Next the solution was dried with vacuum drying, followed by being calcined (650°C, 6 hours). 25.5g of 85% H_3PO_4 was dissolved in 3000mL of distilled water and 200g of calcined sample was added to the solution, followed by stirring 3 hours at the room temperature. Next the solution was dried with vacuum drying, followed by being calcined (650°C, 6 hours).

Comparative Example 2: Preparation of P-Mn/ZSM-5

[0060] 40.5 g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ was dissolved in 3000mL of distilled water, and 200g of ZSM-5 was added to the solution, followed by stirring 3 hours at the room temperature. Next the solution was dried with vacuum drying, followed by being calcined (650°C, 6 hours). 25.5g of 85% H_3PO_4 was dissolved in 3000mL of distilled water and 200g of calcined sample was added to the solution, followed by stirring 3 hours at the room temperature. Next the solution was dried with vacuum drying, followed by being calcined (650°C, 6 hours).

Comparative Example 3: Preparation of P/ZSM-5

[0061] 25.5g of 85% H_3PO_4 was dissolved in 3000mL of distilled water and 200g of ZSM-5 was added to the solution, followed by stirring 3 hours at the room temperature. Next the solution was dried with vacuum drying, followed by being calcined (650°C, 6 hours).

Comparative Example 4~6

[0062] The micro-spherical catalyst for catalytic cracking was prepared with using the sample of comparative example 4-6, with following procedure.

[0063] For preparing the zeolite slurry, 120g sample of comparative example 1 was added to 200g of distilled water, followed by stirring. For preparing the clay slurry, 144g of clay was added to 176g of distilled water, followed by stirring. 439g of alumina sol (solid contents 8.4%, pH2~3) was used for binding zeolite and clay to make the micro-spherical catalyst. Zeolite slurry, clay slurry and alumina sol were stirred homogenously, followed by spraying, and drying. Next, thus prepared material was calcined at 650°C for 6 hours to form the molded catalyst of the comparative example 4. With same procedure and method, molded catalysts of comparative examples 5 and 6 were prepared using zeolite of comparative examples 2 and 3.

Comparative Example 7

[0064] For preparing the zeolite slurry, 120g sample of comparative example 1 was slowly added to 200g of distilled water, followed by stirring. For preparing the clay slurry, 144g of clay was slowly added to 176g of distilled water, followed by stirring. For forming inorganic binder to make the micro-spherical catalyst, 439g of alumina sol (solid contents 8.4%, pH2~3) and 33.1g of 85% H_3PO_4 were homogeneously mixed. Zeolite slurry, clay slurry and alumina sol- H_3PO_4 mixture were stirred homogenously, followed by spraying, drying. Next, this was calcined at 650°C for 6 hours, and formed the molded catalyst of the comparative example 7.

[0065] Chemical compositions of catalyst of comparative example 4-7 are summarized in the below table 1.

Table 1

Catalyst	Comparative example 4	Comparative example 5	Comparative example 6	Comparative example 7
composition, weight %				
Zeolite	29.3	33.9	37.0	27.5
Clay	47.9	47.9	47.9	44.8
Al_2O_3	12.3	12.3	12.3	11.5
SiO_2	-	-	-	-
P_2O_5	2.9	2.9	2.9	9.1
LaO	5.1	-	-	4.7
MnO_2	2.6	3.0	-	2.5

Example 1-2

[0066] 4.5g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ was added to 376mL of distilled water, and 120g of ZSM-5 was added to this solution, followed by stirring at 60°C for 6 hours. Next, using high viscosity slurry mixer, 144g of clay was slowly added to this solution and stirred for 3 hours. For preparing inorganic binder, 439g of alumina sol (solid contents 8.4%, pH2~3), 30.5g of 85% H_3PO_4 and 1.8g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ were mixed at 35°C for 8 hours. Above zeolite-clay slurry, and inorganic binder

were homogeneously mixed, followed by spraying, drying. Next, after being calcined at 650°C for 6 hours, catalyst of example 1 was formed.

[0067] The same procedure of example 1 was performed except for the different amount of samples (11.2g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ was used for ZSM-5 forming, 3.1g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ and 43.8g of H_3PO_4 was used for inorganic binder) to form the catalyst of example 2.

Example 3-4

[0068] 4.5g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ was added to 376mL of distilled water, and 120g of ZSM-5 was added to this solution, followed by stirring at 60°C for 6 hours. Next, using high viscosity slurry mixer, 144g of clay was slowly added to this solution for 3 hours. 56.7g of Pseudo Boehmite (Al_2O_3 contents 72%) was dispersed in 498g of distilled water. Next for preparing inorganic former, this dispersed Pseudo Boehmite solution, 30.5g of 85% H_3PO_4 and 1.8g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ were mixed at 35°C for 8 hours. For preparing the inorganic former, 5.32g of formic acid was added to this mixture and stirred until being stabilized. Above zeolite-clay slurry, and inorganic binder were homogeneously mixed, followed by spraying, drying. Next, after being calcined at 650°C for 6 hours, catalyst of example 3 was formed.

[0069] The same procedure of example 3 was performed except for the different amount of samples (15.5g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ was used for ZSM-5 forming, 4.8g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ and 51.3g of H_3PO_4 was used for inorganic binder) to form the catalyst of example 4.

Example 5-6

[0070] 4.5g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ was added to 376mL of distilled water, and 120g of ZSM-5 was added to this solution, followed by stirring at 60°C for 6 hours. Next, using high viscosity slurry mixer, 144g of clay was slowly added to this solution for 3 hours. 23.6g of water glass (SiO_2 29%) was added to 199g solution of aluminium sulfate (Al_2O_3 8%), and mixed. Next for preparing inorganic binder, this solution, 15.95g of 85% H_3PO_4 and 1.8g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ were mixed at 35°C for 8 hours. Above zeolite-clay slurry and inorganic binder were homogeneously mixed, followed by spraying, drying. Next, after being calcined at 650°C for 6 hours, catalyst of example 5 was formed.

[0071] The same procedure of example 5 was performed except for the different amount of samples (11.4g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ was used for ZSM-5 forming, 5.8g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ and 71.2g of H_3PO_4 was used for inorganic binder) to form the catalyst of example 6.

[0072] Chemical compositions of catalyst of example 1-6 are summarized in the below table 2

Table 2

Catalyst	example 1	example 2	example 3	example 4	example 5	example 6
Composition, weight %						
Zeolite	37.2	35.9	36.7	34.8	40.2	35.5
Clay	44.7	43.1	44.1	41.6	48.1	42.6
Al_2O_3	11.4	11.0	12.5	11.8	5.3	4.7
SiO_2	-	-	-	-	2.3	2.0
P_2O_5	5.8	8.1	5.8	9.2	3.3	13.0
LaO	-	-	-	-	-	-
MnO_2	0.9	1.9	0.9	2.6	0.8	2.2

Example 7-8

[0073] 4.5g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ was added to 376mL of distilled water, and 90g of ZSM-5 was added to this solution, followed by stirring at 60°C for 6 hours. Next, using high viscosity slurry mixer, 144g of clay was slowly added to this solution for 3 hours. 62.4g of $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ was added to 220mL of distilled water, followed that 21.5g of 85% H_3PO_4 and 1.3g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ were mixed at 35°C for 8 hours. Above zeolite-clay slurry and this solution were homogeneously mixed, followed by spraying, drying. Next, after being calcined at 650°C for 6 hours, catalyst of example 7 was formed.

[0074] The same procedure of example 7 was performed except for the different amount of samples (120g of ZSM-5 was used, 11.4g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ was used for ZSM-5 forming, 5.8g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ and 61.2g of H_3PO_4 was used for

inorganic binder) to form the catalyst of example 8.

Comparative example 8

[0075] 120g of ZSM-5 was stirred with 376mL of distilled water, at the room temperature for 6 hours. Next, using high viscosity slurry mixer, 144g of clay was slowly added to this solution for 3 hours. 62.4g of $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ was added to 220mL of distilled water, followed by mixing with 21.5g of 85% H_3PO_4 . Above zeolite-clay slurry and this solution were homogeneously mixed, followed by spraying, drying. Next, after being calcined at 650°C for 6 hours, catalyst of comparative example 8 was formed.

Comparative example 9

[0076] 13.2g of 85% H_3PO_4 was added to the 576mL of aqueous solution (22.8g of $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ and 222.6g of $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$ were dissolved), followed by stirring for 3 hours. This solution was titrated with ammonia water to make pH=11. After removing the sediment, drying at 100°C, and being calcined at 650°C for 5 hours, MnAlPOx was prepared. 32.6g of MnAlPOx and 120g of ZSM-5 were added to 200g of distilled water, and mixed to form $\text{MnAlPOx}/\text{ZSM-5}$ slurry. For preparing the clay slurry, 111.4g of clay and 176g of distilled water were used with above procedure. 439g of alumina sol (solid contents 8.4%, pH 2~3), zeolite slurry and clay slurry were homogeneously mixed, followed by spraying, drying, and being calcined at 650°C for 6 hours to form the catalyst of comparative example 9.

[0077] Chemical compositions of catalyst of example 7-8 and comparative example 8-9 are summarized in the below table 3

Table 3

Catalyst	example 7	example 8	Comparative Example 8	Comparative Example 9
Composition weight%				
Zeolite	34.8	37.8	42.0	39.9
Clay	55.8	45.3	50.4	37.0
Al_2O_3	3.3	2.7	3.0	20.0
SiO_2	-	-	-	-
P_2O_5	5.1	11.9	4.6	1.4
LaO	-	-	-	-
MnO_2	1.0	2.38	-	1.7

[0078] ✕ Evaluation of catalyst activity ✕

[0079] For evaluation of catalyst activity, 14 catalyst samples of above comparative examples 4 to 9 and examples 1 to 8 were steamed at 760°C in an atmosphere of 100% steam for 24 hours. The test conditions for evaluation was that reaction temperature was 675°C, weight hourly space velocity (WHSV) was 8/hr, 6g of catalyst was loaded, and naphtha (Boiling point 30~135°C) was used as reactant. Test results are summarized in Tables 4~6.

[0080] From the result, it is obvious that high reaction conversion and high light olefin yield were obtained by introducing Mn and P to make the micro-spherical catalyst according to the present invention. Mn and P are effective for stabilizing zeolite, inorganic binder and clay. Also Mn and P protect the acid center of zeolite to achieve high light olefin yield.

Table 4

	Comparative example 4	Comparative example 5	Comparative example 6	Comparative example 7	Comparative example 8	Comparative example 9
Product distribution, wt%						
$\text{C}_2=$	16.5	14.4	13.5	13.3	11.2	14.4
$\text{C}_3=$	19.6	19.4	19.4	19.2	18.1	19.2

(continued)

	Comparative example 4	Comparative example 5	Comparative example 6	Comparative example 7	Comparative example 8	Comparative example 9
Product distribution, wt%						
C ₂ -	7.8	7.0	8.5	6.4	5.6	7.5
C ₃ -	4.9	4.3	6.0	3.2	2.3	4.1
C ₄	10.6	10.8	10.2	10.1	9.5	10.7
C ₅ (iso-C ₅ , n-C ₅)	7.4	9.3	10.6	10.6	12.3	6.8

Table 5

	Example 1	Example 2	Example 3	Example 4	Example 5
Product distribution, wt%					
C ₂ =	19.5	19.4	21.8	21.7	20.5
C ₃ =	21.4	20.9	20.0	19.8	21.5
C ₂ -	9.1	9.6	10.1	10.1	9.1
C ₃ -	5.8	5.8	4.5	4.3	3.9
C ₄	9.3	9.7	7.4	7.3	8.4
C ₅ (iso-C ₅ , n-C ₅)	4.4	4.2	1.6	1.4	3.6

Table 6

	Example 6	Example 7	Example 8
Product distribution, wt%			
C ₂ =	20.8	21.7	21.9
C ₃ =	20.3	21.0	22.3
C ₂ -	9.8	9.5	9.8
C ₃ -	4.7	3.7	4.1
C ₄	8.5	7.6	7.0
C ₅ (iso-C ₅ , n-C ₅)	2.0	2.4	2.4

[0081] As disclosed above, the catalyst according to the present invention is characterized in that for achieving high yield of light olefin, zeolite acid point is treated with Mn, and in order for thus treated zeolite to achieve high activity in the catalyst structure, P and Mn are used to stabilize the inorganic oxide binder and matrix component. The present method of preparing catalyst has advantages from the point of expense comparing the prior art which generally comprises complicated embedding steps for zeolite.

Claims

1. A hydrocarbon cracking catalyst for preparing light olefin from C₄ or more than C₄ hydrocarbon having a boiling point of 30 ~ 200 °C, which is **characterized in that** 0.01~5.0 wt% of MnO₂ and 1~15 wt% of P₂O₅ are simultaneously

supported on a catalyst component, wherein the catalyst component comprises 1 ~ 50wt% of zeolite, 21 ~ 70wt% of clay, and 1 ~ 40wt% of an inorganic oxide, wherein zeolite is ZSM-5 zeolite, and wherein both MnO_2 and P_2O_5 are embedded on each of the zeolite, the clay and the inorganic oxide in the catalyst component.

2. The catalyst of claim 1, wherein the zeolite has a Si/Al molar ratio of less than or equal to 200.
3. The catalyst of claim 1, wherein the inorganic oxide is Al_2O_3 , SiO_2 or $\text{Al}_2\text{O}_3\text{-SiO}_2$.
4. A method of preparing the cracking catalyst according to claim 1, the method comprising the steps of:
 - (a) Mixing zeolite, clay and inorganic oxide precursor with phosphorus precursor and manganese precursor with stirring to prepare a mixing slurry; and
 - (b) spray drying the mixing slurry, followed by calcinations.
5. A method of preparing the cracking catalyst according to claim 4, wherein the step (a) comprises the steps of:
 - (i) mixing zeolite with manganese precursor, followed by adding the clay and stirring the mixture to prepare a zeolite/clay slurry;
 - (ii) mixing the inorganic oxide precursor with phosphorus precursor and manganese precursor with stirring to prepare an inorganic oxide slurry; and
 - (iii) mixing the zeolite/clay slurry and the inorganic oxide slurry uniformly.
6. A method of preparing the cracking catalyst according to claim 4, wherein the step (a) comprises the steps of:
 - (i) mixing zeolite with manganese precursor to prepare a zeolite slurry;
 - (ii) mixing the inorganic oxide precursor with phosphorus precursor and manganese precursor with stirring to prepare an inorganic oxide slurry; and
 - (iii) mixing the zeolite slurry, the inorganic oxide slurry, and clay uniformly.
7. A method of preparing the cracking catalyst according to claim 4, wherein the step (a) is carried out by mixing zeolite, clay and inorganic oxide precursor, phosphorus precursor and manganese precursor at the same time with stirring.
8. The method of any of claims 4-7, wherein the zeolite has a Si/Al molar ratio of less than or equal to 200.
9. The method of any of claims 4-8, wherein the inorganic oxide precursor comprises Al_2O_3 , SiO_2 or $\text{Al}_2\text{O}_3\text{-SiO}_2$, and is formed of sol, gel, or solution.
10. The method of any of claims 4-9, wherein the manganese precursor is sulfate, nitrate, chloride or acetate compound of manganese.
11. The method of any of claims 4-10, wherein the phosphorus precursor is aqueous phosphorus compound selected from a group of H_3PO_4 , $(\text{NH}_4)_3\text{PO}_4$, $\text{H}(\text{NH}_4)_2(\text{PO}_4)$ and $\text{H}_2(\text{NH}_4)\text{PO}_4$.

Patentansprüche

1. Kohlenwasserstoff-Crackkatalysator für die Herstellung von Leichtolefin aus C4- oder mehr als C4-Kohlenwasserstoffen, die einen Siedepunkt von 30 ~ 200 °C aufweisen, der **dadurch gekennzeichnet ist, dass** 0,01 ~ 5,0 Gew.-% MnO_2 und 1 ~ 15 Gew.-% P_2O_5 gleichzeitig auf einer Katalysatorkomponente geträgert sind, wobei die Katalysatorkomponente 1 ~ 50 Gew.-% Zeolith, 21 ~ 70 Gew.-% Ton und 1 ~ 40 Gew.-% eines anorganischen Oxids umfasst, wobei der Zeolith ZSM-5-Zeolith ist und wobei sowohl MnO_2 als auch P_2O_5 auf jedem von dem Zeolith, dem Ton und dem anorganischen Oxid in der Katalysatorkomponente eingebettet sind.
2. Katalysator nach Anspruch 1, wobei der Zeolith ein Si/Al-Molverhältnis von weniger als oder gleich 200 aufweist.
3. Katalysator nach Anspruch 1, wobei das anorganische Oxid Al_2O_3 , SiO_2 oder $\text{Al}_2\text{O}_3\text{-SiO}_2$ ist.

4. Verfahren für die Herstellung eines Crackkatalysators nach Anspruch 1, wobei das Verfahren die Schritte umfasst des:

- (a) Mischens von Zeolith, Ton und anorganischem Oxidvorläufer mit Phosphorvorläufer und Manganvorläufer unter Rühren, um eine Mischaufschlämmung herzustellen, und
- (b) Sprühtrocknens der Mischaufschlämmung, gefolgt von Calcinierungen.

5. Verfahren für die Herstellung des Crackkatalysators nach Anspruch 4, wobei der Schritt (a) die Schritte umfasst des:

- (i) Mischens von Zeolith mit Manganvorläufer, gefolgt von Zugabe des Tons und Rühren der Mischung, um eine Zeolith-/Tonaufschlämmung herzustellen,
- (ii) Mischens des anorganischen Oxidvorläufers mit Phosphorvorläufer und Manganvorläufer unter Rühren, um eine anorganische Oxidaufschlämmung herzustellen, und
- (iii) gleichförmigen Mischens der Zeolith-/Tonaufschlämmung und der anorganischen Oxidaufschlämmung.

6. Verfahren für die Herstellung des Crackkatalysators nach Anspruch 4, wobei der Schritt (a) die Schritte umfasst des:

- (i) Mischens von Zeolith mit Manganvorläufer, um eine Zeolithaufschlämmung herzustellen,
- (ii) Mischens des anorganischen Oxidvorläufers mit Phosphorvorläufer und Manganvorläufer unter Rühren, um eine anorganische Oxidaufschlämmung herzustellen, und
- (iii) gleichförmigen Mischens der Zeolithaufschlämmung, der anorganischen Oxidaufschlämmung und von Ton.

7. Verfahren für die Herstellung des Crackkatalysators nach Anspruch 4, wobei der Schritt (a) durch Mischens von Zeolith, Ton und anorganischem Oxidvorläufer, Phosphorvorläufer und Manganvorläufer gleichzeitig unter Rühren ausgeführt wird.

8. Verfahren nach irgendeinem der Ansprüche 4-7, wobei der Zeolith ein Si/Al-Molverhältnis von weniger als oder gleich 200 aufweist.

9. Verfahren nach irgendeinem der Ansprüche 4-8, wobei der anorganische Oxidvorläufer Al_2O_3 , SiO_2 oder $\text{Al}_2\text{O}_3\text{-SiO}_2$ umfasst und aus Sol, Gel oder Lösung gebildet wird.

10. Verfahren nach irgendeinem der Ansprüche 4-9, wobei der Manganvorläufer eine Sulfat-, Nitrat-, Chlorid- oder Acetatverbindung von Mangan ist.

11. Verfahren nach irgendeinem der Ansprüche 4-10, wobei der Phosphorvorläufer eine wässrige Phosphorverbindung ist, die aus einer Gruppe von H_3PO_4 , $(\text{NH}_4)_3\text{PO}_4$, $\text{H}(\text{NH}_4)_2(\text{PO}_4)$ und $\text{H}_2(\text{NH}_4)\text{PO}_4$ ausgewählt ist.

Revendications

1. Catalyseur de craquage d'hydrocarbure pour préparer une oléfine légère à partir d'hydrocarbure en C4 ou supérieur à C4 présentant un point d'ébullition de 30 ~ 200°C, qui est **caractérisé en ce que** 0,01 ~ 5,0 % en masse de MnO_2 et 1 ~ 15 % en masse de P_2O_5 sont simultanément supportés sur un constituant de catalyseur, dans lequel le constituant de catalyseur comprend 1 ~ 50 % en masse de zéolithe, 21 ~ 70 % en masse d'argile, et 1 ~ 40 % en masse d'un oxyde inorganique, dans lequel la zéolithe est une zéolithe ZSM-5, et dans lequel à la fois MnO_2 et P_2O_5 sont incorporés sur chacun de la zéolithe, l'argile et l'oxyde inorganique dans le constituant de catalyseur.

2. Catalyseur selon la revendication 1, dans lequel la zéolithe présente un rapport molaire Si/Al inférieur ou égal à 200.

3. Catalyseur selon la revendication 1, dans lequel l'oxyde inorganique est Al_2O_3 , SiO_2 ou $\text{Al}_2\text{O}_3\text{-SiO}_2$.

4. Procédé de préparation du catalyseur de craquage selon la revendication 1, le procédé comprenant les étapes de :

- (a) mélange de zéolithe, d'argile et de précurseur d'oxyde inorganique avec un précurseur de phosphore et un précurseur de manganèse en agitant pour préparer une suspension de mélange ; et
- (b) séchage par pulvérisation de la suspension de mélange, suivi par des calcinations.

5. Procédé de préparation du catalyseur de craquage selon la revendication 4, dans lequel l'étape (a) comprend les étapes de :

(i) mélange de zéolithe avec un précurseur de manganèse, suivi par l'addition de l'argile et agitation du mélange pour préparer une suspension de zéolithe/argile ;
 (ii) mélange du précurseur d'oxyde inorganique avec un précurseur de phosphore et un précurseur de manganèse en agitant pour préparer une suspension d'oxyde inorganique ; et
 (iii) mélange de la suspension de zéolithe/argile et de la suspension d'oxyde inorganique uniformément.

6. Procédé de préparation du catalyseur de craquage selon la revendication 4, dans lequel l'étape (a) comprend les étapes de :

(i) mélange de zéolithe avec un précurseur de manganèse pour préparer une suspension de zéolithe ;
 (ii) mélange du précurseur d'oxyde inorganique avec un précurseur de phosphore et un précurseur de manganèse en agitant pour préparer une suspension d'oxyde inorganique ; et
 (iii) mélange de la suspension de zéolithe, la suspension d'oxyde inorganique, et l'argile uniformément.

7. Procédé de préparation du catalyseur de craquage selon la revendication 4, dans lequel l'étape (a) est réalisée en mélangeant la zéolithe, l'argile et un précurseur d'oxyde inorganique, un précurseur de phosphore et un précurseur de manganèse simultanément en agitant.

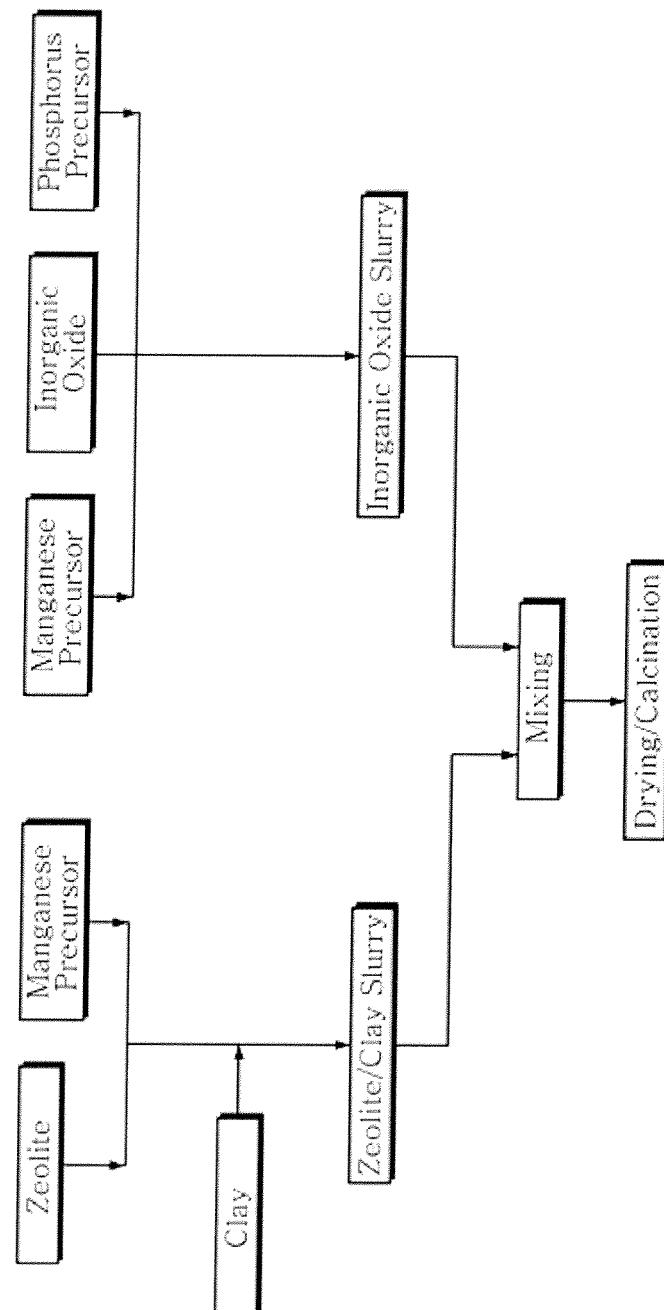
8. Procédé selon l'une quelconque des revendications 4-7, dans lequel la zéolithe présente un rapport molaire Si/Al inférieur ou égal à 200.

9. Procédé selon l'une quelconque des revendications 4-8, dans lequel le précurseur d'oxyde inorganique comprend Al_2O_3 , SiO_2 ou $\text{Al}_2\text{O}_3\text{-SiO}_2$, et est formé de sol, gel, ou solution.

10. Procédé selon l'une quelconque des revendications 4-9, dans lequel le précurseur de manganèse est un composé de sulfate, nitrate, chlorure ou acétate de manganèse.

11. Procédé selon l'une quelconque des revendications 4-10, dans lequel le précurseur de phosphore est un composé aqueux de phosphore choisi dans un groupe formé de H_3PO_4 , $(\text{NH}_4)_3\text{PO}_4$, $\text{H}(\text{NH}_4)_2(\text{PO})_4$ et $\text{H}_2(\text{NH}_4)\text{PO}_4$.

FIG. 1



REFERENCES CITED IN THE DESCRIPTION

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